

Periodic Trends Section Review Answers

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6.3 Periodic Trends. Prentice Hall Chemistry, 2005. STUDY. PLAY. Atomic radii generally _____ as you move from left to right in a period. (energy levels, increases, charge, decreases) decreases.

6.3 Periodic Trends Flashcards | Quizlet

Answers for Comparing Atomic Sizes Here are answers for the questions above. a. Li, C, F All are in the same period and thus have the same number of energy levels. Therefore, the important factor is the nuclear charge. Li is the largest because it has the smallest nuclear charge and pulls the ... Microsoft Word - Periodic Trends Worksheet Author:

Periodic Trends Worksheet - sd308.org

Concept Review: Trends in the Periodic Table 1. ionization energy 2. bond radius 3. electron affinity 4. electronegativity 5. increases, decreasing 6. d 7. a 8. c 9. The electron cloud model is based on the probability of finding an electron at a specific location. As you move far-ther out from the nucleus, the proba-bility of finding electrons ...

Skills Worksheet Concept Review

• Interpret group trends in atomic radii, ionic radii, ionization energies, and ... Answer the following in the space provided. 20. For the following pairs of atoms, tell which one of each pair has the largest ionic radius. a. ... 14.2 Periodic Trends Section Review

14.2 Periodic Trends Section Review - LPS

Section 6.3 Periodic Trends 171 Group Trends in Atomic Size In the Figure 6.14 graph, atomic radius is plotted versus atomic number. Look at the data for the alkali metals and noble gases. The atomic radius within these groups increases as the atomic number increases. This increase is an example of a trend.

6.3 Periodic Trends - schoolwires.henry.k12.ga.us

Periodic Trends >Trends in Ionization Energy The energy required to remove an electron from an atom is called ionization energy. •The energy required to remove the first electron from an atom is called the first ionization energy. •The energy required to remove an electron from an ion with a 1+ charge is called the second ionization energy. 6.3

6.3 Periodic Trends - Mr. Walk

Review of Periodic Trends You will need to periodic table to complete this activity. Choose the correct answer for each question. Show all questions <= => Generally speaking, the group of elements with the highest first ionization energy is: ? Group 1 ? Group 16 ? ...

Review of Periodic Trends - sciencegeek.net

174 Chapter 6 • The Periodic Table and Periodic Law Section 66.1.1 Development of the Modern Periodic Table MAIN Idea The periodic table evolved over time as scientists discovered more useful ways to compare and organize the elements.

Chapter 6: The Periodic Table and Periodic Law

Section 6.3 Periodic Trends 3 sessions 11/ 2 blocks P LS 1. Trace the development and identify key features of the periodic table. 2. Explain why elements in the same group have similar properties. 3. Identify the four blocks of the periodic table based on electron configuration. 4. Compare period and group trends of several properties. 5.

Chapter 6: The Periodic Table and Periodic Law

Periodic Trends Worksheet (Answer Key 1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:)) Cu) K) Ni) Br))) Largest - K Smallest - Br) Explain why you made these choices.) Atomic radius decreases as you go left to right across a period. Potassium is in the far

Periodic Trends Worksheet Answers

SECTION 6.3 PERIODIC TRENDS . 1. Explain why a magnesium atom is smaller than atoms of both sodium and calcium. 2. Predict the size of the astatine (At) atom compared to that of tellurium (Te). Explain your prediction. 3. Would you expect a Cl⁻ ion to be larger or smaller than an Mg²⁺ ion? Explain. 4.

SECTION 6.3 PERIODIC TRENDS - SharpSchool

Section 6.3 Periodic Trends 1. c 2. c 3. d 4. b 5. a 6. a 7. Ionization energy is the energy required to remove an electron from a gaseous atom. 8. A high ionization-energy value indicates that the atom has a strong hold on its electrons and is not likely to lose an outer electron and form a positive ion.

Ch 6 Study Guide answers

Periodic Trends > Trends in Ionization Energy Group and Periodic Trends in Ionization Energy First ionization energy tends to decrease from top to bottom within a group and increase from left to right across a period. 6.3

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Periodic Table Review Directions: Complete the chart using a periodic table. ... Directions: label the trends of the periodic table by adding arrows and descriptions. & Page 2 of 2 9.1 Use a periodic table of the elements to answer these questions. 1. The following graphics represent the nuclei of atoms. Using a periodic table of elements, fill ...

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